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Metallurgy about Worksheet الملف

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HIGHER SECONDARY SECOND YEAR CHEMISTRY

Choose the correct answer:

- Bauxite has the composition
a) Al_2O_3 b) $\text{Al}_2\text{O}_3 \cdot n\text{H}_2\text{O}$ c) $\text{Fe}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ d) None of these
- Roasting of sulphide ore gives the gas (A). (A) is a colourless gas. Aqueous solution of (A) is acidic. The gas (A) is
a) CO_2 b) SO_3 c) SO_2 d) H_2S
- Which one of the following reaction represents calcinations?
a) $2\text{Zn} + \text{O}_2 \longrightarrow 2\text{ZnO}$ b) $2\text{ZnS} + 3\text{O}_2 \longrightarrow 2\text{ZnO} + 2\text{SO}_2$
c) $\text{MgCO}_3 \longrightarrow \text{MgO} + \text{CO}_2$ d) Both (a) and (c)
- The metal oxide which cannot be reduced to metal by carbon is
a) PbO b) Al_2O_3 c) ZnO d) FeO
- Which of the metal is extracted by Hall-Heroult process?
a) Al b) Ni c) Cu d) Zn
- Which of the following statements, about the advantage of roasting of sulphide ore before reduction is not true?
a) ΔG_f° of sulphide is greater than those for CS_2 and H_2S .
b) ΔG_r° is negative for roasting of sulphide ore to oxide
c) Roasting of the sulphide to its oxide is thermodynamically feasible.
d) Carbon and hydrogen are suitable reducing agents for metal sulphides.
- Match items in column - I with the items of column - II and assign the correct code.

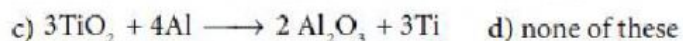
a) $\left(\frac{\Delta S^0}{\Delta T}\right)$ is negative

b) $\left(\frac{\Delta G^0}{\Delta T}\right)$ is positive

c) $\left(\frac{\Delta G^0}{\Delta T}\right)$ is negative

d) initially $\left(\frac{\Delta T}{\Delta G^0}\right)$ is positive, after 700°C ,
 $\left(\frac{\Delta G^0}{\Delta T}\right)$ is negative

23. Which of the following reduction is not thermodynamically feasible?



24. Which of the following is not true with respect to Ellingham diagram?

a) Free energy changes follow a straight line. Deviation occurs when there is a phase change.

b) The graph for the formation of CO_2 is a straight line almost parallel to free energy axis.

c) Negative slope of CO shows that it becomes more stable with increase in temperature.

d) Positive slope of metal oxides shows that their stabilities decrease with increase in temperature.