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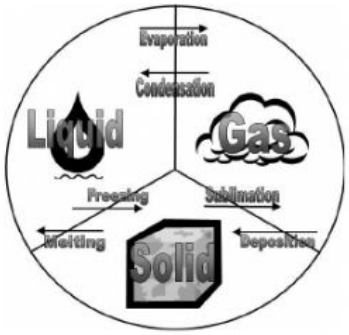
Changes of State

Chapter 13 Study Guide-Part 4



Lesson Objectives

- ❑ Identify the conditions necessary for sublimation.
- ❑ Describe how equilibrium conditions are represented in a phase diagram.

True-False: Classify each of these statements as always true (AT); sometimes true (ST); or never true (NT).

_____ 1. Water could be made to boil at 105°C by increasing the pressure.

_____ 2. At 101.3 kPa, the normal boiling point and melting point of water are the same.

_____ 3. Water has more than one triple point.

_____ 4. The sublimation point of a substance refers to the temperature and pressure at which the substance exists in all three phases of matter.

_____ 5. A phase diagram gives information on changes in mass of solids, liquids, and gases.

_____ 6. Below the triple point for water, decreasing the pressure will not change water vapor to ice.

Vocabulary: Match each description with the correct term.

_____ 7. The change of a solid to the liquid state.

_____ 8. Defines the triple point of water.

_____ 9. A method of removing water from food, using sublimation.

A. freeze drying

B. melting

C. phase diagram

_____ 10. Normal boiling point of water.	D. sublimation
_____ 11. Graph that shows the relationship among the states of a substance.	E. 0.016°C, 0.61 kPa
_____ 12. The change of a solid to a vapor without passing through the liquid phase.	F. 100°C at 101.3 kPa

Drag and Drop: Drag the appropriate term to the correct statement.

carbon dioxide **equilibrium** **phase** **triple point** **sublimation**

vapor pressure **0.61 kPa** **0.016°C**

The change that occurs when a solid goes directly to the gas or vapor state without first becoming a liquid is (13) _____. This change can occur because solids, like liquids, have a (14) _____. Substances that sublime include iodine and solid (15)_____(dry ice).

A graph that shows the relationship between the states of a substance is called (16)_____ diagram. On this diagram, a line between two phases shows the conditions at which the phases are in (17) _____. The (18) _____ is the only set of conditions at which solid, liquid, and gas phases coexist. The triple point for water is a temperature of (19) _____ and a pressure of (20) _____.