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More files for 9th Grade , Subject Chemistry , Term 1

Chemistry Worksheet	1
Worksheet about Hydrogen	2
Mole to Grams and Grams to Moles Conversions Worksheet	3
Worksheet about Chemistry	4
Chemistry Worksheet	5
Worksheet about fundamentals of organic reactions	6
Worksheet about Chemistry Pharmaceutics	7

12. A compound formed by elements X and Y crystallizes in a cubic structure in which the X atoms are at the corners of a cube and the Y atoms are at the face-centres. The formula of the compound is
 (a) XY_3 (b) X_3Y (c) XY (d) XY_2
13. A compound is formed by elements A and B. The crystalline cubic structure has the A atoms at the corners of the cube and B atoms at the body centre. The simplest formula of the compound is
 (a) AB (b) A_6B (c) AB_6 (d) A_8B_4
14. If NaCl is doped with 10^{-4} mol % of SrCl_2 , the concentration of cation vacancies will be
 ($N_A = 6.02 \times 10^{23} \text{mol}^{-1}$)
 (a) $6.02 \times 10^{16} \text{mol}^{-1}$ (b) $6.02 \times 10^{17} \text{mol}^{-1}$
 (c) $6.02 \times 10^{14} \text{mol}^{-1}$ (d) $6.02 \times 10^{15} \text{mol}^{-1}$
15. The appearance of colour in solid alkali metal halides is generally due to
 (a) Schottky defect (b) Frenkel defect
 (c) Interstitial positions (d) F-centres
16. Schottky defect in crystal is observed when
 (a) an ion leaves its normal site and occupies an interstitial site.
 (b) unequal number of cations and anions are missing from the lattice
 (c) density of the crystal is increased.
 (d) equal number of cations and anions are missing from the lattice.
17. When electrons are trapped into the crystal in anion vacancy, the defect is known as
 (a) Schottky defect (b) Frenkel defect
 (c) Stoichiometric defect (d) F-centres
18. Which of the following is not a characteristic of a crystalline solid?
 a) Definite and characteristic heat of fusion
 b) Isotropic nature
 c) A regular periodically repeated pattern of arrangement of constituent particles in the entire crystal
 d) A true solid
19. Iodine molecules are held in the crystal lattice by
 a) London forces b) dipole-dipole interactions
 c) Covalent bonds d) coulombic forces
20. Which one of the following is a covalent crystal?
 a) Rock b) Ice c) Quartz d) Dry ice