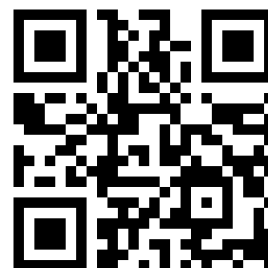


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الملف Worksheet about Molecular Kinetic Theory Part 3 Study Guide

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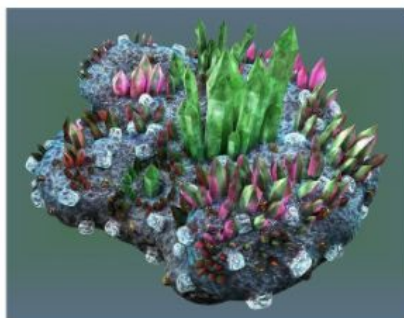
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The Nature of Solids

Chapter 13 Study Guide-Part 3

Lesson Objectives

- ❑ Evaluate how the way particles are organized explains the properties of solids.
- ❑ Identify the factors that determine the shape of a crystal solid.
- ❑ Explain how allotropes of an element are different.



Vocabulary Matching:

- | | |
|--|------------------|
| _____ 1. Describes a solid in which the particles are randomly arranged. | A. amorphous |
| _____ 2. The temperature at which a solid changes into a liquid. into a liquid. | B. crystal |
| _____ 3. One of the seven crystal systems. | C. glasses |
| _____ 4. Has a regular 3-Dimensional arrangement of particles. | D. melting point |
| _____ 5. Dense state of matter with a fixed shape and not easily compressed. | E. rhombohedral |
| _____ 6. Transparent fusion products of inorganic substances that has cooled to a rigid state without crystallization. | F. solid |
| _____ 7. The smallest group of particles within a crystal that retains geometric shape of the crystal | G. unit cell |

Completion: Drag and Drop

amorphous

compress

crystalline

fixed

high

freezing point

melts

melting point

lattice

unit cell

Solids tend to be dense and difficult to (8) _____. They do not flow or take the shape of their containers, like liquids do, because the particles in solids vibrate around (9)_____ points. When a solid is heated until its particles vibrate so rapidly that they are no longer held in fixed positions, the solid (10) _____. The (11) _____ is the temperature at which a solid changes to a liquid. The melting and (12) _____ of a substance are at the same temperature. In general, ionic solids tend to have relatively (13)_____ melting points, while molecular solids tend to have relatively low melting points. Most solids are (14) _____. The particles are arranged in a pattern known as crystal (15) _____. The smallest subunit of a crystal lattice is the (16) _____. Some solids lack an ordered internal structure and are called (17) _____ solids.

True-False: Classify the following as always true(AT); sometimes true (ST); or never true (NT).

- _____ 18. When the atoms in a solid have a random arrangement, the solid is a glass.
- _____ 19. Glasses do not melt at a definite temperature, but soften gradually.
- _____ 20. Allotropes are two or more different elements that exist in the same state with the same crystal system.
- _____ 21. Solid substances can exist in more than one form.
- _____ 22. The type of bonding that exists between the atoms in a crystal tends to determine the melting point of the solid.