تغطية اختبار نهاية الفصل

تم تحميل هذا الملف من موقع المناهج الإماراتية	
موقع المناهج ⇦ المناهج الإماراتية ⇦ الصف التاسع ⇦ علوم ⇦ الفصل الأول ⇦ ملفات متنوعة ⇦ الملف	
تاريخ إضافة الملف على موقع المناهج: 2026-11-15 22:56:49	
المزيد من مادة علوم:	ملفات ا كتب للمعلم ا كتب للطالب ا اختبارات الكترونية ا اختبارات ا حلول ا عروض بوربوينت ا أوراق عمل منهج انجليزي ا ملخصات وتقارير ا مذكرات وبنوك ا الامتحان النهائي ا للمدرس

التواصل الاجتماعي بحسب الصف التاسع					
صفحة المناهج الإماراتية على فيسبوك					

المزيد من الملفات بحسب الصف التاسع والمادة علوم في الفصل الأول	
مذكرة مزودة بأسئلة اختبار من متعدد وأمثلة محلولة على جميع المواضيع	1
انسباير عام - organisms and their relationships	2
أنسباير متقدم - chemical reactions	3
ملزمة تحتوي على شرح دروس الفصل كامل بالإضافة الى تمارين بعد كل درس 2026-2027	4
أسئلة على درس المادة - انسباير	5

تحميل تطبيق المناهج للهواتف الذكية	
تطبيق المناهج على Store App (آبل)	تطبيق المناهج على Play Google (أندرويد)

End of Term 1 Exam Coverage

Q1	BIO.3.1.02.022 Identify examples of chemical reactions that support main functions of living organisms where reactants and products rearrange to form ATP, ADP and inorganic phosphate	Figure 13	11

Reactants and Products :

Chemical reaction: is the process by which atoms or groups of atoms in substances are reorganized into different substances.

▶ **Rust is a result of a chemical change called a chemical reaction.**

Chemical reaction has taken place include production of

- heat or light
- formation of a new gas, liquid, or solid.
- Change color



Figure 13 After a chemical change, such as rusting, a new substance is formed. During a physical change, such as water boiling, the chemical makeup of the water is not altered.

☐ Chemical bonds are broken and formed during chemical reactions.

☐ The rust on the train in Figure 13 is a compound called iron oxide.

☐ (Fe_2O_3) formed when oxygen (O_2) in the air reacted with iron (Fe).

Physical change : alters a substance's appearance, but not its composition.

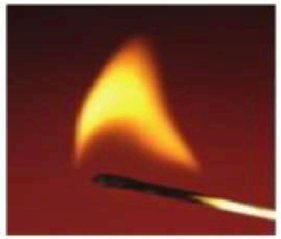
Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q1	BIO.3.1.02.022 Identify examples of chemical reactions that support main functions of living organisms where reactants and products rearrange to form ATP, ADP and inorganic phosphate	Figure 13	11

1

What does the image shows?

A. A covalent bond.
B. A physical property.
C. Chemical reaction.
D. Van der Waals forces.



2

Which observation can provide evidence that a chemical reaction has occurred?

A. A change in shape only.	**B.** The production of heat or light.	**C.** A change from solid to liquid.	**D.** A change in appearance without a change in composition.

3

Which of the following is an example of a physical change?

A. Iron rusting.	**B.** Formation of a new substance.	**C.** Water changing from liquid to gas.	**D.** Chemical bonds being broken and formed.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q2	BIO.3.1.02.022 Identify examples of chemical reactions that support main functions of living organisms where reactants and products rearrange to form ATP, ADP and inorganic phosphate	Figure 14	13

Energy of Reaction

Activation energy : is the minimum amount of energy needed for reactants to form products in a chemical reaction

- Some reactions rarely happen because they have a very high activation energy.
- The peak in the graph represents the activation energy, the amount of energy that must be added to the system to make the reaction occur.

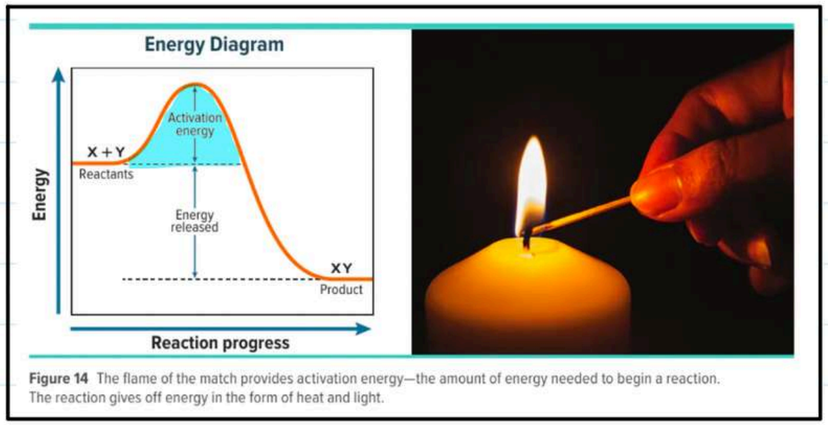


Figure 14 The flame of the match provides activation energy—the amount of energy needed to begin a reaction. The reaction gives off energy in the form of heat and light.

For example, you know that a candle will not burn until you light its wick. The flame provides the activation energy for the reaction that occurs between the substances in the candle wick and the oxygen in the air. In this example, once the reaction begins, no further input of energy is needed, and the candle continues to burn on its own

X + Y

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q2	BIO.3.1.02.022 Identify examples of chemical reactions that support main functions of living organisms where reactants and products rearrange to form ATP, ADP and inorganic phosphate	Figure 14	13

1

The minimum amount of energy needed for a chemical reaction is called

A.	Principle mass	**B.**	Activation energy	**C.**	Product threshold	**D.**	Absorbent energy

2

Why does a candle require a flame before it begins to burn?

A.	The flame provides the activation energy needed to start the reaction	**B.**	The flame provides the products.	**C.**	The flame removes oxygen from the reaction.	**D.**	The flame acts as a reactant.

3

In an energy diagram, what does the peak of the graph represent?

A.	The products.	**B.**	The reactants.	**C.**	The activation energy.	**D.**	The amount of products formed.



Q3	BIO.3.1.02.021: Identify examples of reactions catalyzed by enzymes that occur in living systems, describing their importance in living organism and explaining why an enzyme deficiency results in the inability to perform a specific function in life.	Figure 17	15

Enzymes

All living things are driven by chemical reactions.

- Additional substances are needed to reduce activation energy and reaction time in living organisms.

Catalyst : is a substance that lowers the activation energy needed to start a chemical reaction.

- Catalysts do not change the amount of product produced, nor are they used up during the reaction.
- Special proteins called **enzymes** are the biological catalysts that speed up the rate of chemical reactions in biological processes.

Enzymes are essential to many life processes.

Most enzymes are specific to one reaction. Like , amylase

Lower activation energy = speed up the rate of chemical reactions = Reduce reaction time

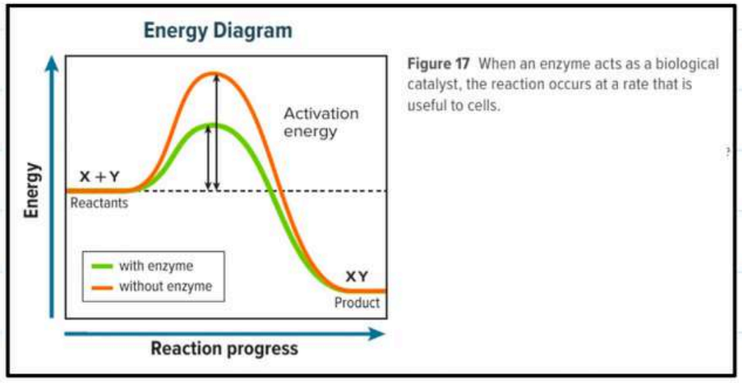


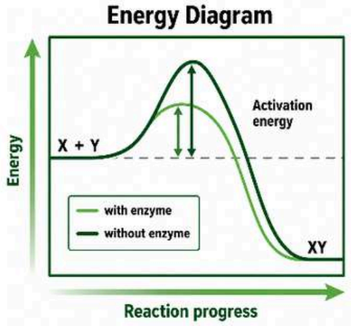
Figure 17 When an enzyme acts as a biological catalyst, the reaction occurs at a rate that is useful to cells.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q3	BIO.3.1.02.021: Identify examples of reactions catalyzed by enzymes that occur in living systems, describing their importance in living organism and explaining why an enzyme deficiency results in the inability to perform a specific function in life.	Figure 17	15

1

How can enzymes benefit chemical reactions, as shown in this energy diagram?



A. lowering activation energy	B. raising activation energy	C. slowing reactions	D. increasing the amount of product

2

Which of the following is a substance that lowers the activation energy?

A. A substrate.	**B.** A catalyst.	**C.** A reactant.	**D.** An anion.

3

Why are enzymes essential to many life processes?

A. They stop chemical reactions in cells.	**B.** They lower the activation energy of biochemical reactions, allowing them to occur quickly.	**C.** They create matter in cells.	**D.** They replace all reactants in biochemical reactions.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q4	BIO.3.1.02.021: Identify examples of reactions catalyzed by enzymes that occur in living systems, describing their importance in living organism and explaining why an enzyme deficiency results in the inability to perform a specific function in life	Figure 18	16

Enzymes- Substrate Reaction : The biological Process

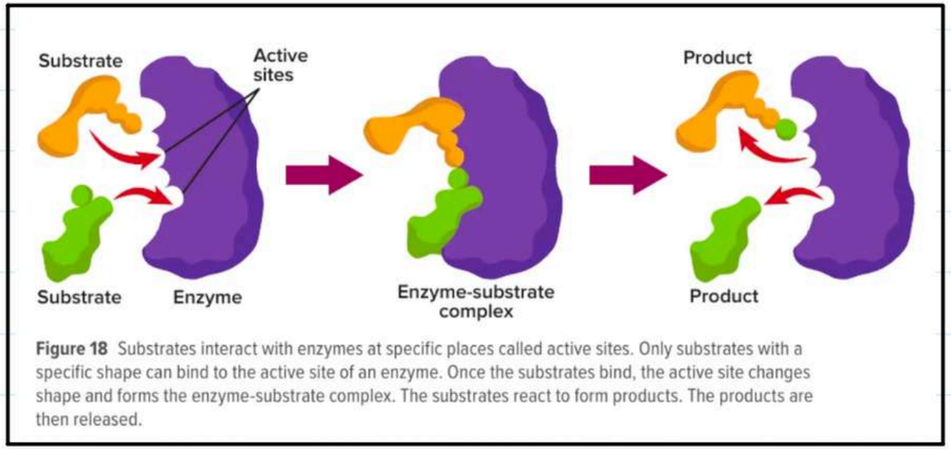


Figure 18 Substrates interact with enzymes at specific places called active sites. Only substrates with a specific shape can bind to the active site of an enzyme. Once the substrates bind, the active site changes shape and forms the enzyme-substrate complex. The substrates react to form products. The products are then released.

The reactants that bind to an enzyme are called **substrates**.

The specific location where a substrate binds on an enzyme is called **the active site.**

- Factors such as pH, temperature, and other substances affect enzyme activity.

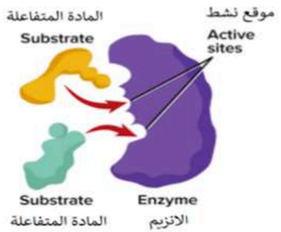
- Only substrates with the same size and shape as the active site will bind to the enzyme.
- Once the substrates bind to the active site, the active site changes shape and forms the enzyme-substrate complex.
- The enzyme-substrate complex helps chemical bonds in the reactants to be broken and new bonds to form-the substrates react to form products.
- Photosynthesis and cellular respiration provide energy for the cell with the help of enzymes.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q4	BIO.3.1.02.021: Identify examples of reactions catalyzed by enzymes that occur in living systems, describing their importance in living organism and explaining why an enzyme deficiency results in the inability to perform a specific function in life	Figure 18	16

1

Why is the active site of an enzyme important to the enzyme activity?



A. It raises the activation energy of a reaction	**B.** It allows the enzyme to interact with a large variety of substrates	**C.** It allows the enzyme to catalyze very specific reactions	**D.** It allows interaction to continue without stopping

2

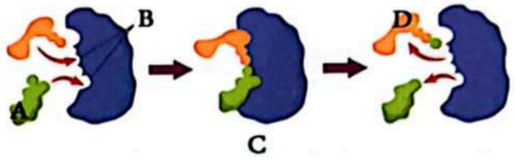
What is the function of the enzyme-substrate complex?

A. destroy the enzyme	**B.** Stop chemical reaction	**C.** Break and forming chemical bond	**D.** Prevent bondage at the active site

3

The figure below represents the enzyme action.

Which of the following letters indicates the product?

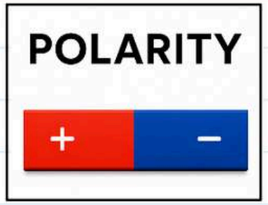


A	B	C	D

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q5	BIO.3.1.02.020 Identify the solubility as the ability of a solute to dissolve in a solvent and explain the properties of water that make it the universal solvent	Figure 19	18

Intermolecular Forces :

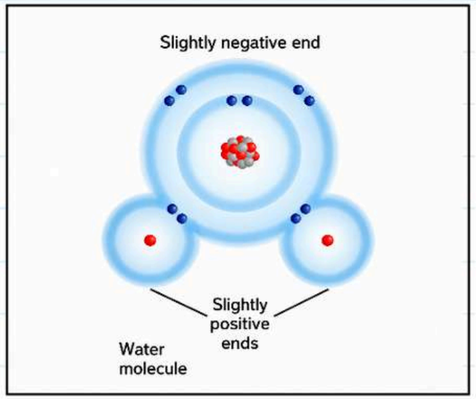


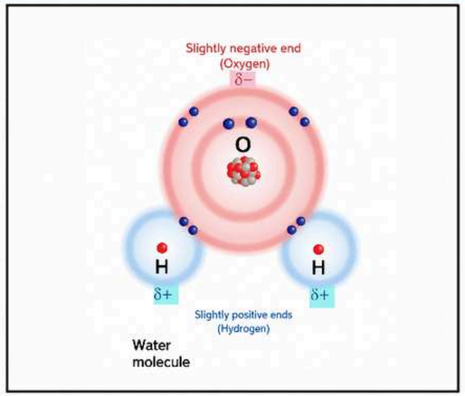
Polarity: is the property of having two opposite poles, or ends.

For example, a magnet has polarity, with a north pole and a south pole

Polar molecules: are molecules that have an unequal distribution of charges. meaning that they have oppositely charged regions.

For example, water (H_2O) is a polar molecule.





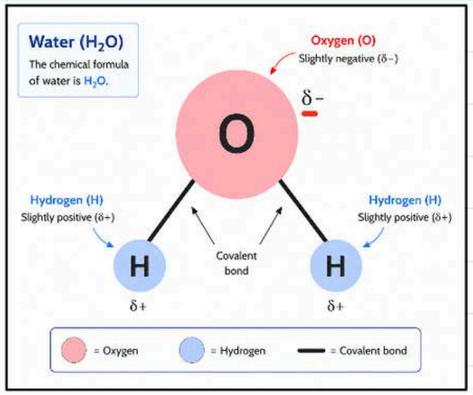
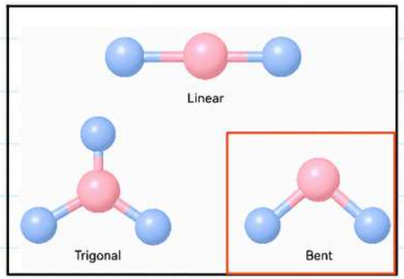


Figure 19 Water molecules have a bent shape, and electrons are not shared equally between hydrogen and oxygen.

This results in a molecule with oppositely charged regions.

Unequal distribution of electrons ⟶ Unequal distribution of charge ⟶ Polar molecule.



Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q5	BIO.3.1.02.020 Identify the solubility as the ability of a solute to dissolve in a solvent and explain the properties of water that make it the universal solvent	Figure 19	18

1

What is a polar molecule?

A. a molecule with oppositely charged regions.	**B.** a molecule that produces hydroxide ions (OH-) in solution.	**C.** a large molecule formed by joining smaller molecules together.	**D.** a molecule that has lost or gained electrons.

2

Which statement best describes the polarity of a water molecule?

A. It has only a positive charge.	**B.** It has only a negative charge.	**C.** It has an equal distribution of electrical charges.	**D.** It has a slightly negative end and slightly positive ends.

3

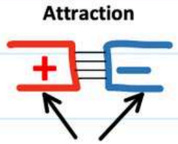
Why is water considered a polar molecule?

A. Its electrons are not shared equally, and the molecule has a bent shape.	**B.** it contains only hydrogen atoms.	**C.** Its electrons are shared equally between hydrogen and oxygen.	**D.** Its atoms have equal attraction for electrons.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q6	BIO.3.1.02.020: Identify the solubility as the ability of a solute to dissolve in a solvent and explain the properties of water that make it the universal solvent	Figure 21	19

Van der Waals forces :



Van der Waals forces: Are weak electrostatic attractions between molecules caused by regions of opposite charge.

Van der Waals forces are weaker than covalent and ionic bonds, but they are important in biological processes.

- Van der Waals forces are easier to break apart, making them ideal for short-term interaction.
- **Strong bonds** ⟶ covalent & ionic
- **Weak attractions** ⟶ van der Waals force



Figure 20. Scientists have determined that geckos can climb smooth surfaces because of van der Waals forces between the hairlike structures on their toes and the surface they are climbing.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Intermolecular forces: are electrostatic attractions (attractive force) between different molecules.

- ☐ A compound's chemical properties are mainly determined by its chemical structure
- ☐ A compound's physical characteristics, such as boiling point and melting point are determined by its intermolecular force.
- ☐ The strength of electrostatic attraction depends on :

- The size of the molecule
- The shape of the molecule
- The ability of atom to attract the electrons.

- ☐ The areas of slight positive and negative charges around the water molecule are attracted to the opposite charge of other nearby water molecule.

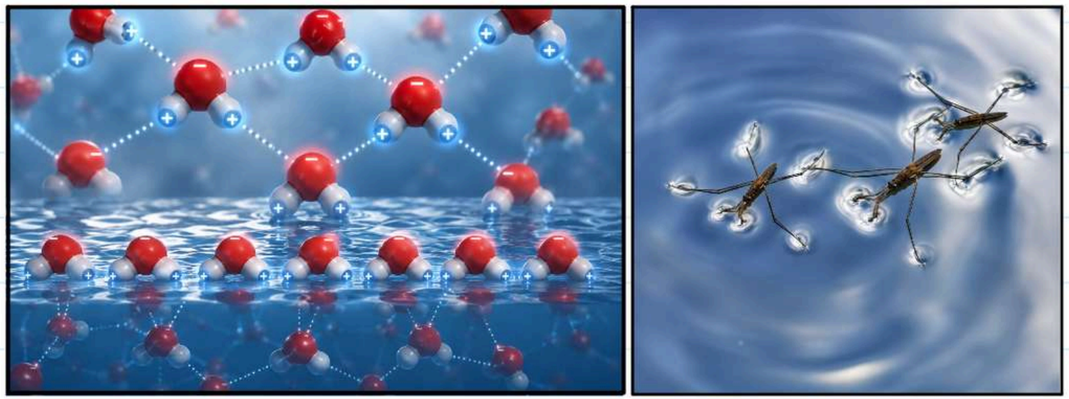


Figure 21. Due to the attractive forces between water molecules, the surface of water supports a water strider.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q6	BIO.3.1.02.020: Identify the solubility as the ability of a solute to dissolve in a solvent and explain the properties of water that make it the universal solvent	Figure 21	19

1

How are water molecules in a rain drop held together?

A. Attraction between positive charges	**B.** Attraction between negative charges	**C.** Van der Waals forces	**D.** Ionic bonds

2

Why are intermolecular forces, such as van der Waals forces, important to biological function?

A. Intermolecular forces do not play a major role in biological functions.	**B.** They are stronger than ionic or covalent bonds.	**C.** They are easy to form and break apart, making them ideal for short-term interactions.	**D.** They are easy to form and break apart, making them ideal for long-term interactions.

3

Which type of intermolecular force is described as a weak electrostatic attraction between regions of opposite charge in nearby molecules?

A. van der Waals force	**B.** Nuclear force	**C.** Ionic bond	**D.** Covalent bond.

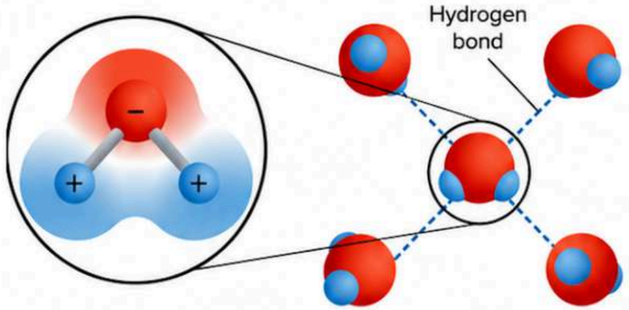
Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q7	BIO.3.1.02.020: Identify the solubility as the ability of a solute to dissolve in a solvent and explain the properties of water that make it the universal solvent	Figure 23	21

Hydrogen Bonding

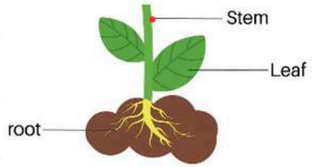
A water molecule is made up of one oxygen atom and two hydrogen atoms. A water molecule is polar in nature. Its bent shape results in a slightly positive charge on the hydrogen atoms and a slightly negative charge on the oxygen atom. As a result, it forms hydrogen bonds.

Water is called the universal solvent because many substances dissolve in it.



Adhesion

Water is adhesive-it forms hydrogen bonds with molecules on other surfaces. **Capillary action is the result of adhesion;** it's why water can travel up the stems of plants. Water travels up the paper, and carries the dye to the middle glass via adhesion.





Capillary action in glass tubes



Cups demonstrating adhesion

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Cohesion

Water is cohesive-the molecules are attracted to each other because of hydrogen bonds. The positive polarity of the hydrogen atom is attracted to the negative polarity of the oxygen atoms. This attraction creates surface tension, which causes water to form droplets and allows insects and leaves to rest on the surface of water.



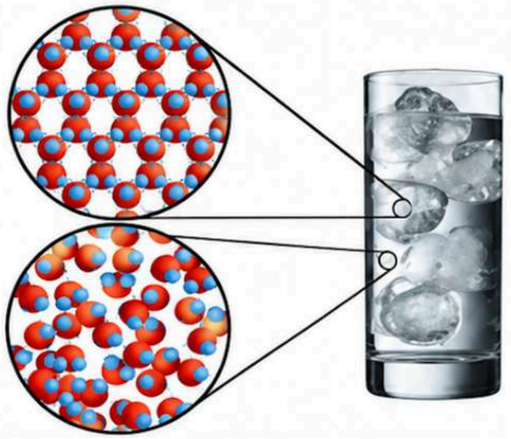
Water strider



Dew on grass

Ice Floats

Liquid water becomes more dense as it cools to 4°C. Yet, ice is less dense than liquid water. As a result, nutrients in bodies of water mix because of changes in water density during spring and fall. Also, fish can survive in winter because ice floats-they continue to live and function in the water beneath the ice.



Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q7	BIO.3.1.02.020: Identify the solubility as the ability of a solute to dissolve in a solvent and explain the properties of water that make it the universal solvent	Figure 23	21

1

Why does water form droplets and support some insects on its surface?

A.	B.	C.	D.
Cohesion creates surface tension through attraction between water molecules.	Adhesion creates ionic bonds between water molecules.	Water loses its polarity at the surface.	Water molecules stop moving at the surface.

2

Which statement correctly explains why ice floats on liquid water?

A.	B.	C.	D.
Liquid water becomes less dense as it cools below 4°C.	Ice contains no water molecules.	Ice is denser than liquid water.	Ice is less dense than liquid water.

3

A scientist observes that water travels upward through the stem of a plant. Which combination of properties of water best explains this movement?

A.	B.	C.	D.
Surface tension and colloid formation	Polarity and heterogeneous mixing	Cohesion and suspension	Adhesion and capillary action

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q8	BIO.3.1.02.018 Distinguish between homogeneous mixtures (solutions) and heterogeneous mixtures	Figure 24	22

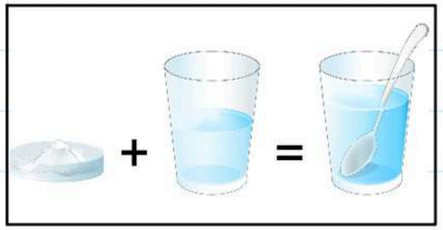
Mixtures with Water

Mixture : is a combination of two or more substances in which each substance retains its individual characteristics and properties.

Homogeneous mixture

Homogeneous: is a mixture with a uniform composition.

- A **solution** is another name for a **homogeneous mixture.**
- Air is a solution of gases.
- In a solution, there are two components: a **solvent** and a **solute**.



solvent: is a substance in which another substance is dissolved.

solute: is the substance that is dissolved in the solvent.

In the drink mix, **water is the solvent**, and the **powder is the solute.**



Figure 24 The drink mix forms a homogeneous mixture in water. The particles of the solute (drink mix) are dissolved and spread throughout the solvent (water).

- A mixture of salt and water is another example of a solution because the solute (**salt**) dissolves completely in the solvent (**water**).
- Saliva moistens your mouth and begins the digestion of some foods. It is a solution that contains water, proteins, and salts.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Heterogeneous mixture

In a **heterogeneous mixture**, the components remain distinct; you can tell what they are individually.

A salad containing lettuce, croutons, and dressing is a **heterogeneous mixture**

- Compare the mixture of sand and water to the solution of salt and water

▶ Sand and water form a type of **heterogeneous mixture** called a **suspension**.

- Over time, the particles in a suspension settle to the bottom.

A **colloid** is a **heterogeneous mixture** in which the particles do not settle out.

▶ You are probably familiar with many colloids, including smoke, butter, milk, paint, and ink.

▶ Blood, is a colloid made up of plasma, cells, and other substances.

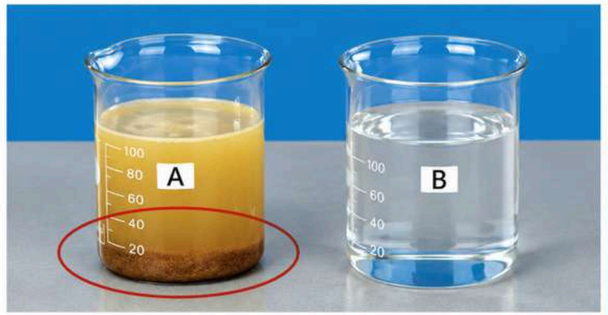


Figure 25 Sand and water form a heterogeneous mixture; you can see both the liquid and the solid. The homogeneous mixture of salt and water is a liquid; you cannot see the salt.



Figure 26 Blood is a heterogeneous mixture called a colloid.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q8	BIO.3.1.02.018 Distinguish between homogeneous mixtures (solutions) and heterogeneous mixtures	Figure 24	22

1

What is a homogeneous mixture?

A.	A mixture whose particles always settle to the bottom.
B.	A mixture in which no substances dissolve.
C.	A mixture with a uniform composition throughout.
D.	A mixture in which the components remain visibly distinct.

2

How do solutes and solvents interact?

A.	solvents change the properties of solutes
B.	solvents dissolve in solutes
C.	solutes change the properties of solvents
D.	solutes dissolve in solvents

3

Which example is a homogeneous mixture?

A.	Salt and water	**B.**	Sand and water	**C.**	Salad	**D.**	Croutons and dressing.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q9	BIO.3.1.02.023: Use acid-base indicators or pH test strips to classify solutions as acidic, basic, or neutral.	Figure 28	24

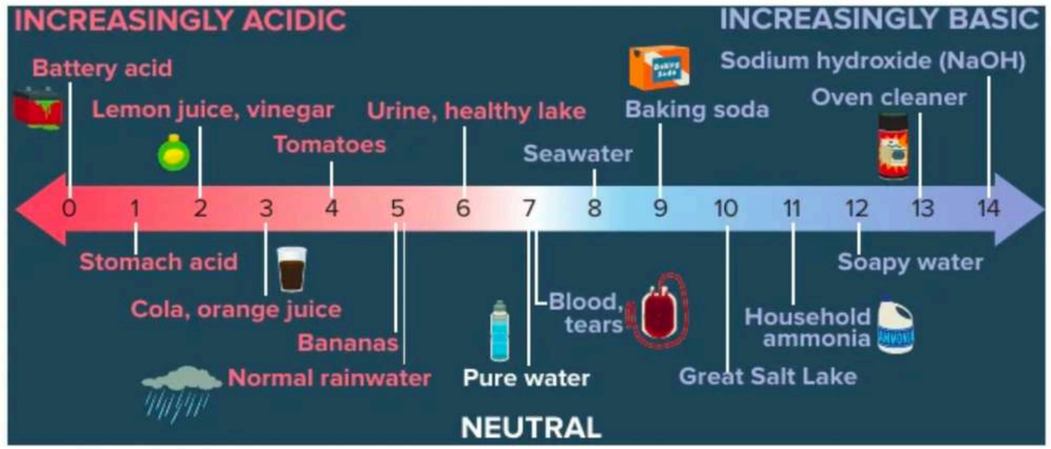


Figure 28 The pH scale is used to indicate the relative strength of acids and bases.

pH and buffers

The measure of concentration of H^+ in a solution is called **pH.**

- ☐ The pH scale is used to indicate the relative strength of acids and bases.
- pure water is neutral; it has a pH value of 7.0
- Acidic solutions have an abundance of (H^+) and have pH values lower than 7
- Basic solutions have more (OH^-)than (H^+) and have pH values higher than 7

Buffer : is mixture that can react with acids or bases to keep the pH within a particular range.

- In cells, buffers keep the pH in a cell within the 6.5 to 7.5 pH range.
- Most cellular processes occur between pH 6.5 and 7.5 to maintain homeostasis.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q9	BIO.3.1.02.023: Use acid-base indicators or pH test strips to classify solutions as acidic, basic, or neutral.	Figure 28	24

1

Which statement is true based on this scale?

Hint : higher concentration of hydrogen ions (H^+) = More acidic



A. Seawater has a higher concentration of hydrogen than tomatoes.	**B.** Tomatoes have a higher concentration of hydrogen than seawater.	**C.** Blood has no hydrogen ions.	**D.** Household ammonia is neutral.

2

A solution has a pH of 4. Which statement is scientifically correct?

A. It is a buffer because its pH is below 7	**B.** It is basic and has an abundance of OH^- ions.	**C.** It is acidic and has a greater concentration of H^+ than a neutral solution.	**D.** It is neutral and has equal amounts of H^+ and OH^- .

3

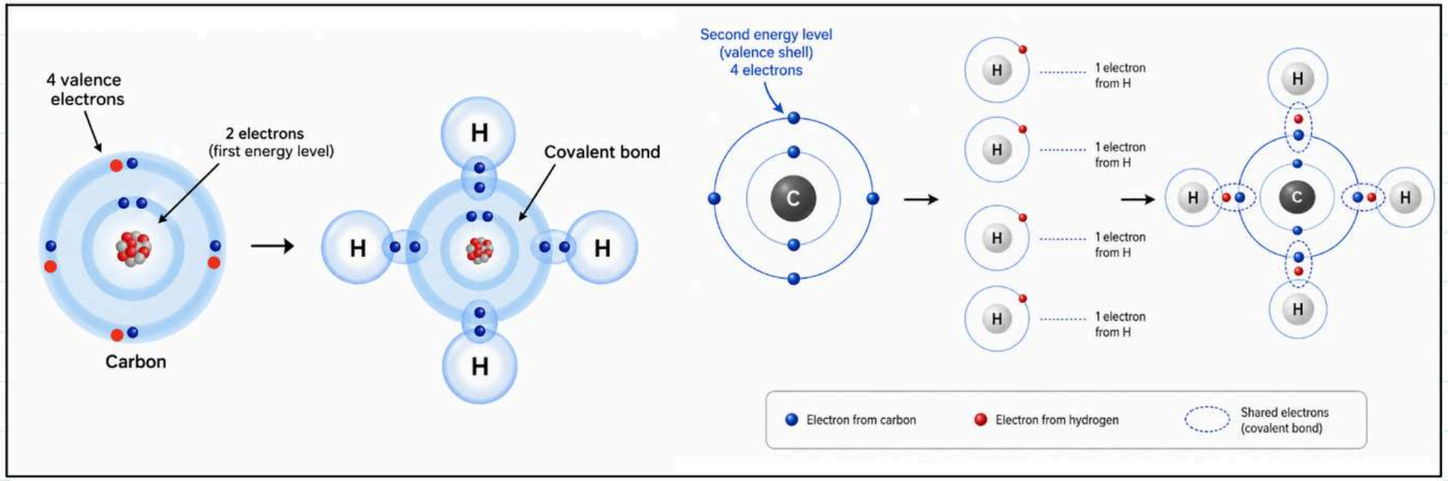
A substance releases H^+ ions when dissolved in water. Which statement correctly describes this substance?

A. It is a solute that cannot affect pH.	**B.** It is an acid, and releasing more H^+ makes the solution more acidic.	**C.** It is a base, and increasing its concentration decreases acidity.	**D.** It is a buffer, and it always has a pH of 7.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Organic Chemistry :

- The element carbon is a component of almost all biological molecules.
- For this reason, life on Earth often is considered carbon-based.
- Because carbon is an essential element, scientists have devoted an entire **branch of chemistry**, called **organic chemistry**.
- to study the organic compounds, which are those compounds containing carbon.
- **carbon has four electrons in its outermost energy level.**
- **Recall :** that the second energy level can hold eight electrons, so **one carbon atom can form four covalent bonds** with other atoms.
- these covalent bonds enable the carbon atoms to bond to each other .



- which results in a variety of important organic compounds. These compounds can be in the shape of **straight chains, branched chains, and rings.**

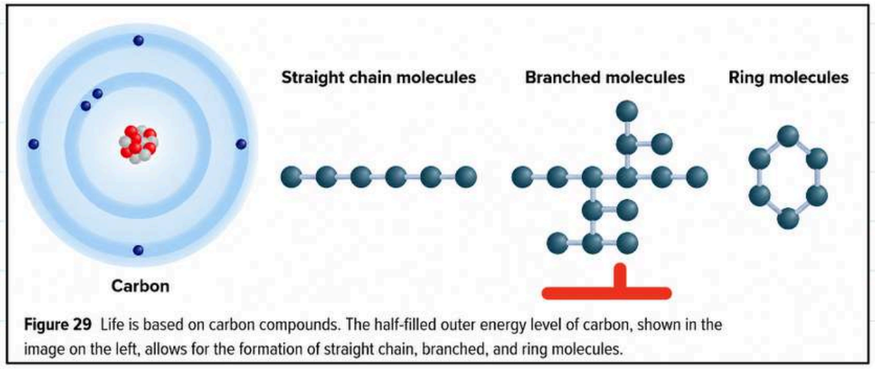


Figure 29 Life is based on carbon compounds. The half-filled outer energy level of carbon, shown in the image on the left, allows for the formation of straight chain, branched, and ring molecules.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos

Q10	BIO.3.1.02.019: Construct and revise an explanation based on evidence for how carbon, hydrogen, and oxygen from sugar molecules may combine with other elements to form amino acids	Figure 29	25

1

Why is life on Earth often considered carbon-based?

A. Carbon is found only in living organisms.	**B.** Carbon is a component of almost all biological molecules.	**C.** Carbon can form only one type of bond.	**D.** Carbon is the most abundant element on Earth.

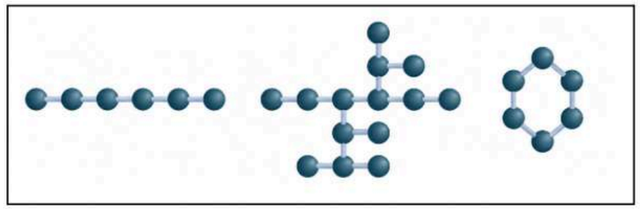
2

What enables carbon atoms to bond to each other and form a variety of important organic compounds?

A. Hydrogen bond	**B.** Ionic bond	**C.** Covalent bond	**D.** Metallic bond

3

Based on Figure, which statement correctly describes the structures formed by carbon atoms?



A. Carbon atoms can form only straight chains.	**B.** Carbon atoms can form straight chains, branched chains, and rings.	**C.** Carbon atoms can form only rings.	**D.** Carbon atoms cannot bond to other carbon atoms.

Grade 9 Biology
@Grade9BiologyUAE · 300 subscribers · 7 videos
Grade 9 Biology | أحياء الصف التاسع ...more
t.me/Grade_9_Biology
Customize channel
Manage videos