

مراجعة المنهج خطة تغطية أسئلة الامتحان النهائي 2026م



تم تحميل هذا الملف من موقع المناهج الإماراتية

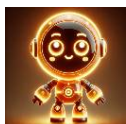
موقع المناهج ⇨ المناهج الإماراتية ⇨ الصف الثاني عشر المتقدم ⇨ كيمياء ⇨ الفصل الثالث ⇨ ملفات متنوعة ⇨ الملف

تاريخ إضافة الملف على موقع المناهج: 09:38:38 2026-05-22

ملفات اكتب للمعلم اكتب للطالب ا اختبارات الكترونية ا اختبارات ا حلول ا عروض بوربوينت ا أوراق عمل
منهج انجليزي ا ملخصات وتقارير ا مذكرات وبنوك ا الامتحان النهائي ا للمدرس

المزيد من مادة
كيمياء:

التواصل الاجتماعي بحسب الصف الثاني عشر المتقدم



صفحة المناهج
الإماراتية على
فيسبوك

الرياضيات

اللغة الانجليزية

اللغة العربية

التربية الاسلامية

المواد على تلغرام

المزيد من الملفات بحسب الصف الثاني عشر المتقدم والمادة كيمياء في الفصل الثالث

الهيكل الوزاري بريدج 2026

1

مقرر الوحدات والدروس المطلوبة في الفصل الثالث منهجي بريدج وانسباير

2

أوراق عمل الوحدة السادسة Electrochemistry منهج انسباير

3

مقرر الوحدات والدروس المطلوبة في الفصل الثالث منهج انسباير Inspire

4

ملخص درس متشكلات الهيدروكربونات من وحدة الهيدروكربونات

5

Academic Year	2025/2026
العام الدراسي	
Term	3
المصطلح	
PLAN C101	
Subject	Chemistry /كيمياء
المادة	
Grade	12
الصف	
Stream	Advanced
المتخصص	
Number of MCQ	30
عدد الأسئلة الموضوعية	
Marks of MCQ	from 2 to 4
درجة الأسئلة الموضوعية	
Number of FRQ	0
عدد الأسئلة المقالية	
Marks per FRQ	0
الدرجات للأسئلة المقالية	
Type of All Questions	MCQ/ الأسئلة الموضوعية
نوع الأسئلة	
Maximum Overall Grade	100
الدرجة القصوى الممكنة	
Exam Duration	120 minutes
مدة الامتحان	
Mode of implementation	Self-Assess
طريقة التطبيق	
Calculator	Allowed
المحاسبة	مسموحة

Question*	Learning Outcomes/Performance Criteria**	Reference(s) in the Student Book (English Version& Arabic Version)	
		Example/Exercise	Page
السؤال*	نتائج التعلم/معايير الأداء**	مثال/تمرين	الصفحة
1	CHM 5.3.05.007.02 Identify components of a voltaic or galvanic cell (anode, cathode, salt bridge or porous barrier, wires, electrolyte compartments); while explaining the role of each component, when does the reaction start and determining the direction of electron and current flow	Textbook + figures 1, 2	25, 26
2	CHM 5.3.05.007.02 Identify components of a voltaic or galvanic cell (anode, cathode, salt bridge or porous barrier, wires, electrolyte compartments); while explaining the role of each component, when does the reaction start and determining the direction of electron and current flow	Textbook + figure 2	26,
3	CHM 5.3.05.007.05 Use the half-cell standard reduction potentials to calculate the electrochemical cell standard potential, while determining whether the redox reactions are spontaneous or non-spontaneous	Textbook + Example problem + Practice problems- check your progress	31, 32, 33, 34, 35
4	CHM 5.3.05.007.04 Write the cell notation and the overall chemical equation for a redox reaction occurring in a voltaic cell	Textbook + Example problem + Practice problems	31, 32, 33
5	CHM 5.3.05.008.02 Use the standard reduction potentials to predict if a reaction occurs or not, while identifying the strongest reducing or oxidizing agent and the substance that is easily oxidized or reduced	Textbook + figure 6	31
6	CHM 5.3.05.012.03 Compare between electrolytic cell and voltaic cell in terms of identifying where will reduction and oxidation processes take place, anode, cathode, direction of electron flow and current flow and spontaneity of the reaction occurring	Textbook + figure 18	47
7	CHM 5.3.05.015.05 Define electroplating while describing how it works, identifying anode, cathode and electrolyte needed for an electrolytic cell in which a selected metal is to be plated on an object (cup or spoon, etc)	Textbook + figure 22	51, 52
8	CHM 5.6.01.001.07 Explain why carbon forms many compounds	Textbook + figure 2	60
9	CHM 5.6.01.012.08 Use different types of models (ball-and-stick model, space-filling model) and formulae (molecular formula, structural formula) to represent a hydrocarbon	Textbook + figure 4	61
10	CHM 5.6.01.001.09 Explain, using Lewis diagram and ball-stick model, the difference between saturated and unsaturated hydrocarbons	Textbook + figure 5	61
11	CHM 5.6.01.013.02 Use IUPAC system to name aliphatic alkanes (straight chain, branched, non-substituted and substituted)	Textbook + tables 1, 2, 3 + Example problem + Practice problems	66, 67, 68, 69, 70, 71
12	CHM 5.6.01.013.03 Use IUPAC system to name cycloalkanes (non branched, branched, non-substituted and substituted)	Textbook + figure 10 + Example problem 2 + Practice problems	71, 72, 73
13	State physical properties of alkanes and its uses	Textbook + figure 11 + table 4	74
14	CHM 5.6.01.013.08 Draw the structure of an alkane given its IUPAC name	Textbook + table 5	76
15	CHM 5.6.01.013.06 Use IUPAC system to name alkenes (straight chain, branched, non-substituted and substituted)	Textbook + figures 12, 13 + Example problem 3 + Practice problems	77, 78, 79
16	CHM 5.6.01.013.07 Use IUPAC system to name alkyne (straight chain and branched)	Textbook + table 6 + check your progress	80, 81
17	CHM 5.6.01.004.01 Define the term structural isomers (and illustrate their similarities and differences)	Textbook + figure 7	82
18	CHM 5.6.01.011.02 Describe the difference between cis- and trans- isomers in terms of geometrical arrangements	Textbook + figures 18, 19	83
19	CHM 5.6.01.003.26 Describe the structure of benzene and its reactivity	Textbook + figure 25	87, 88
20	CHM 5.6.01.009.04 Distinguish among aliphatic and aromatic hydrocarbons	Textbook	88
21	CHM 5.6.01.003.18 Define functional group and give examples	Textbook + table 1	98, 99
22	CHM 5.6.01.013.12 Write the IUPAC name of alkyl halides and aryl halides	Textbook + figures 3, 4 + practice problems	100, 101
23	CHM 5.6.01.013.13 Draw the structure if an alkyl halide/aryl halide given its name	Textbook + figure 4 + practice problems	101
24	CHM 5.6.01.005.06 Define substitution reaction while writing the equation for the substitution reaction including halogenation of alkanes, alkyl halide-alcohol reaction and alkyl halide-amine reaction	Textbook + table 6	109, 110
25	CHM 5.6.01.013.14 Use the IUPAC system to name alcohols	Textbook + table 3 + figures 7, 8	104, 106
26	CHM 5.6.01.013.18 Draw the structure of an ether given its name	Textbook + table 4	107
27	CHM 5.6.01.006.03 Compare and contrast properties of ether and alcohol of similar size and mass (volatility, boiling point, solubility in water)	Textbook	105, 107
28	CHM 5.6.01.013.24 Use the IUPAC system to name carboxylic acids	Textbook + table 9	114
29	CHM 5.6.01.013.26 Use the IUPAC system to name ketones	Textbook + table 8	113
30	CHM 5.6.01.016.06 Explain the physical properties of aldehydes and ketones (polarity, boiling point and solubility)	Textbook	112
* Questions might appear in a different order in the actual exam			
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** As it appears in the textbook (IASE Edition Grade 11 - 12 Advanced - Inquire - Student Edition - volume 2), ILM, and (Main_IP).			
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